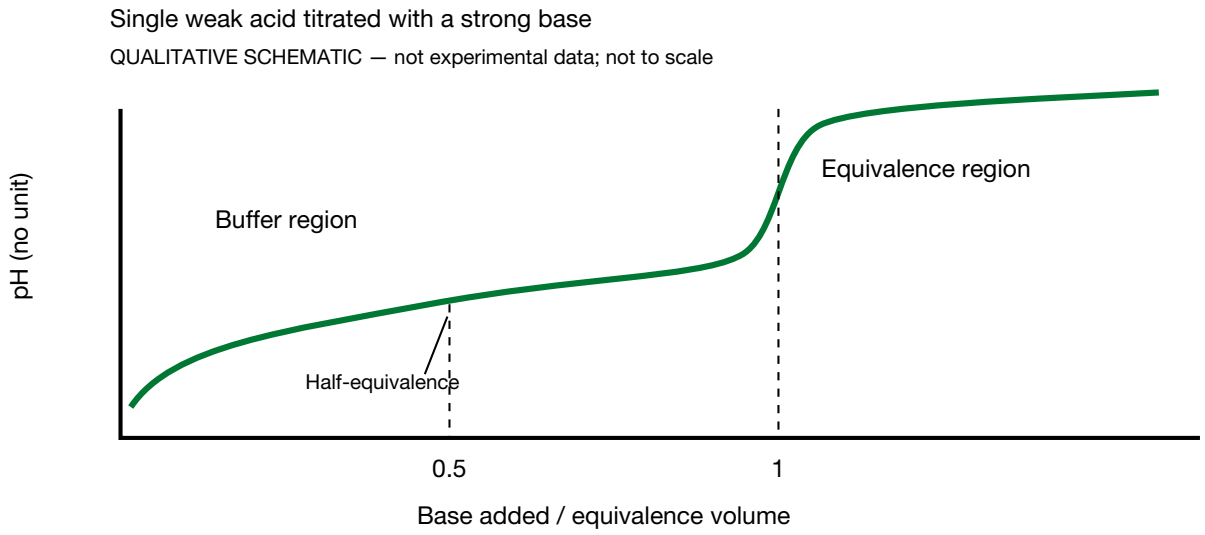
TITRATION CURVE: READ, THEN REASON.

Name the chemical system and locate the regions before calculating. The stem's printed rules still come first.



1 / Name the model

Here: one weak acid and strong base. A polyprotic acid can have multiple equivalence regions; do not transfer this one-step sketch blindly.

2 / Read the axes

The x-axis is a volume ratio, not mL. pH is logarithmic and has no unit; check every label in the actual figure.

3 / Locate the regions

Before equivalence, acid and conjugate base form a buffer. At half-equivalence, pH = pKa for this model.

4 / Use the printed evidence

Locate equivalence rather than assuming it occurs at pH 7. For a weak acid with strong base, it is above 7 under the standard aqueous model.

Method: mapped chemistry page. Chemistry corroboration: OpenStax Chemistry 2e, 14.7.